



Dober dan učenci 9. razredov!

Še nekaj časa nam je ostalo, zato kar nadaljujmo z računanjem.



Pripravite in

Uspešno delo,

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Cilj: reševati naloge.

Naloga: Najprej preveri rešitve nalog prejšnje ure. Nato prepisi ostalo in reši dano nalogo.

Rešitev – Mr.

- a) $Mr(\text{H}_2\text{SO}_4) = 2 \times \text{H} + 1 \times \text{S} + 4 \times \text{O} = 2 \times 1 + 1 \times 32 + 4 \times 16 = 98$
- b) $Mr(\text{Ba}(\text{OH})_2) = 1 \times \text{Ba} + 2 \times \text{O} + 2 \times \text{H} = 1 \times 137 + 2 \times 16 + 2 \times 1 = 171$
- c) $Mr(\text{C}_{15}\text{H}_{31}\text{COOH}) = 16 \times \text{C} + 32 \times \text{H} + 2 \times \text{O} = 16 \times 12 + 32 \times 1 + 2 \times 16 = 256$

Rešitev tabele.

- $Mr(\text{CuCl}_2) = 1 \times \text{Cu} + 2 \times \text{Cl} = 1 \times 64 + 2 \times 36 = 136$
 $M(\text{CuCl}_2) = 136 \text{ g/mol}$
- $Mr(\text{KMnO}_4) = 1 \times \text{K} + 1 \times \text{Mn} + 4 \times \text{O} = 1 \times 39 + 1 \times 55 + 4 \times 16 = 158$
 $M(\text{KMnO}_4) = 158 \text{ g/mol}$
- $Mr(\text{CH}_3\text{CHO}) = 2 \times \text{C} + 4 \times \text{H} + 1 \times \text{O} = 2 \times 12 + 4 \times 1 + 1 \times 16 = 44$
 $M(\text{CH}_3\text{CHO}) = 44 \text{ g/mol}$
- $Mr((\text{NH}_4)_3\text{PO}_4) = 3 \times \text{N} + 12 \times \text{H} + 1 \times \text{P} + 4 \times \text{O} = 3 \times 14 + 12 \times 1 + 1 \times 31 + 4 \times 16 = 149$
 $M((\text{NH}_4)_3\text{PO}_4) = 149 \text{ g/mol}$

Nadaljuješ z novo snovjo → zapis v zvezek.

4. Množina snovi; oznaka = n , enota = mol

- 1 mol je vedno relativna molekulska masa (molska masa) neke snovi.
- Primer:
 $Mr(\text{H}_2\text{O}) = 18 \rightarrow M(\text{H}_2\text{O}) = 18 \text{ g/mol} = 1 \text{ mol}$

